



Process for Converting Kerogen into Petroleum

by Thermo-Chemical Means



Dwight Zwick
President

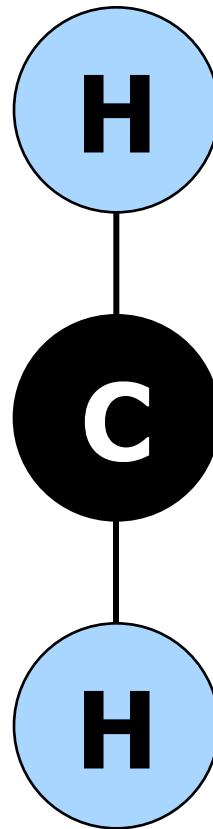
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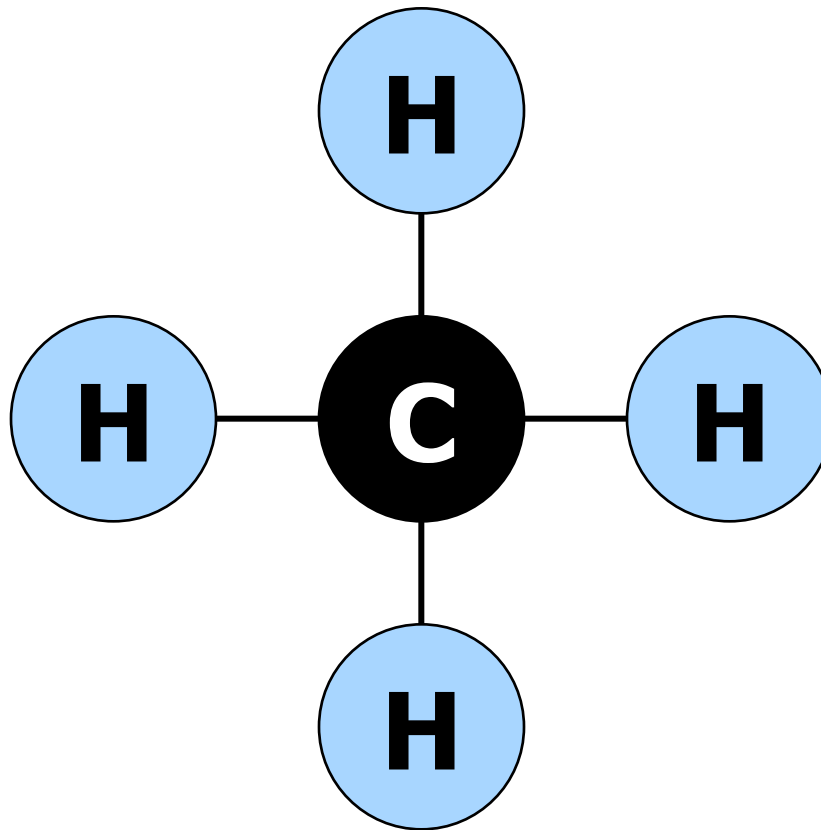
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Unstable Molecule Basic Building Block



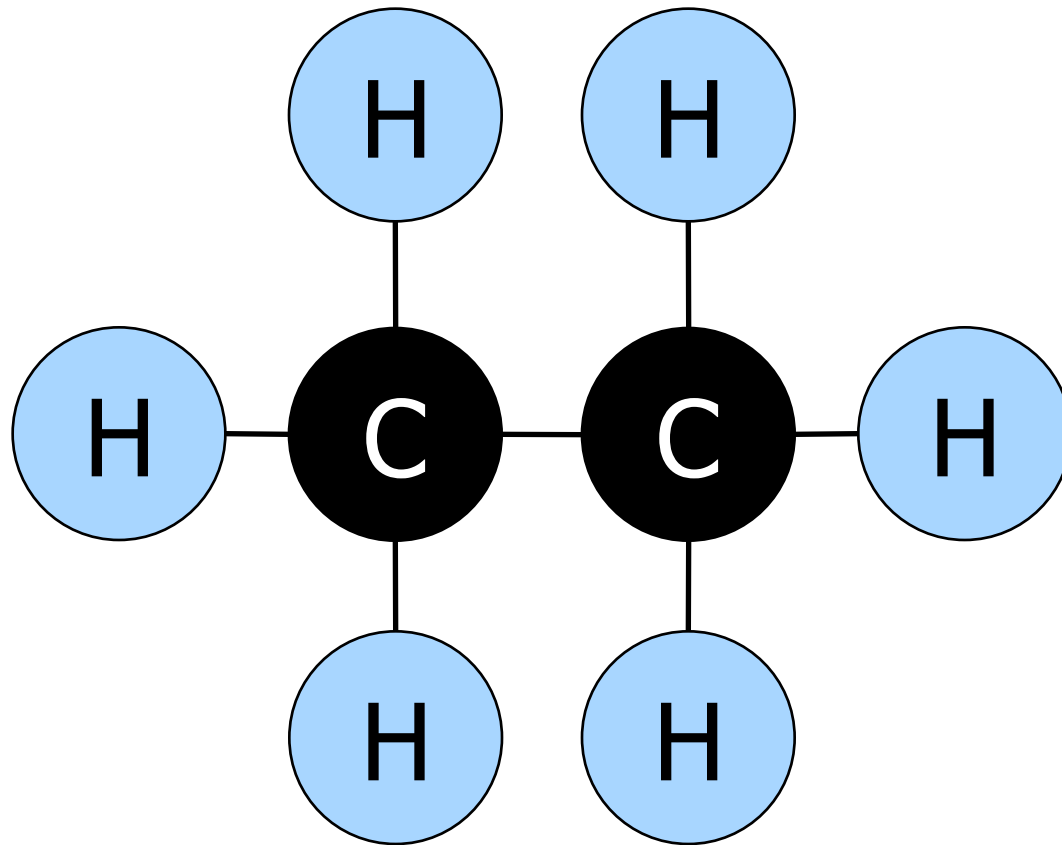


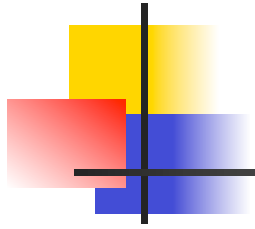
Methane CH₄



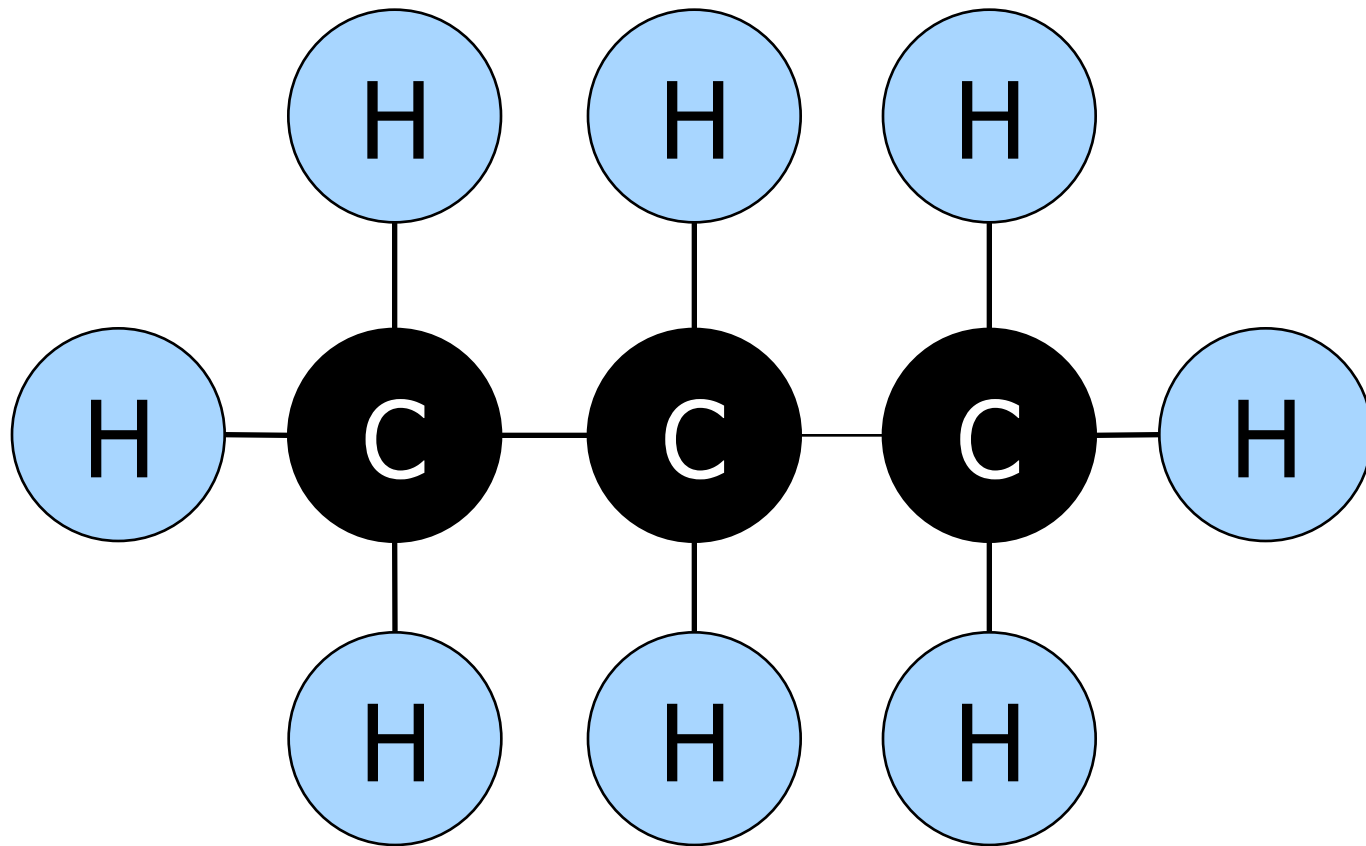


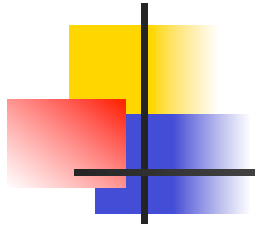
Ethane C_2H_6



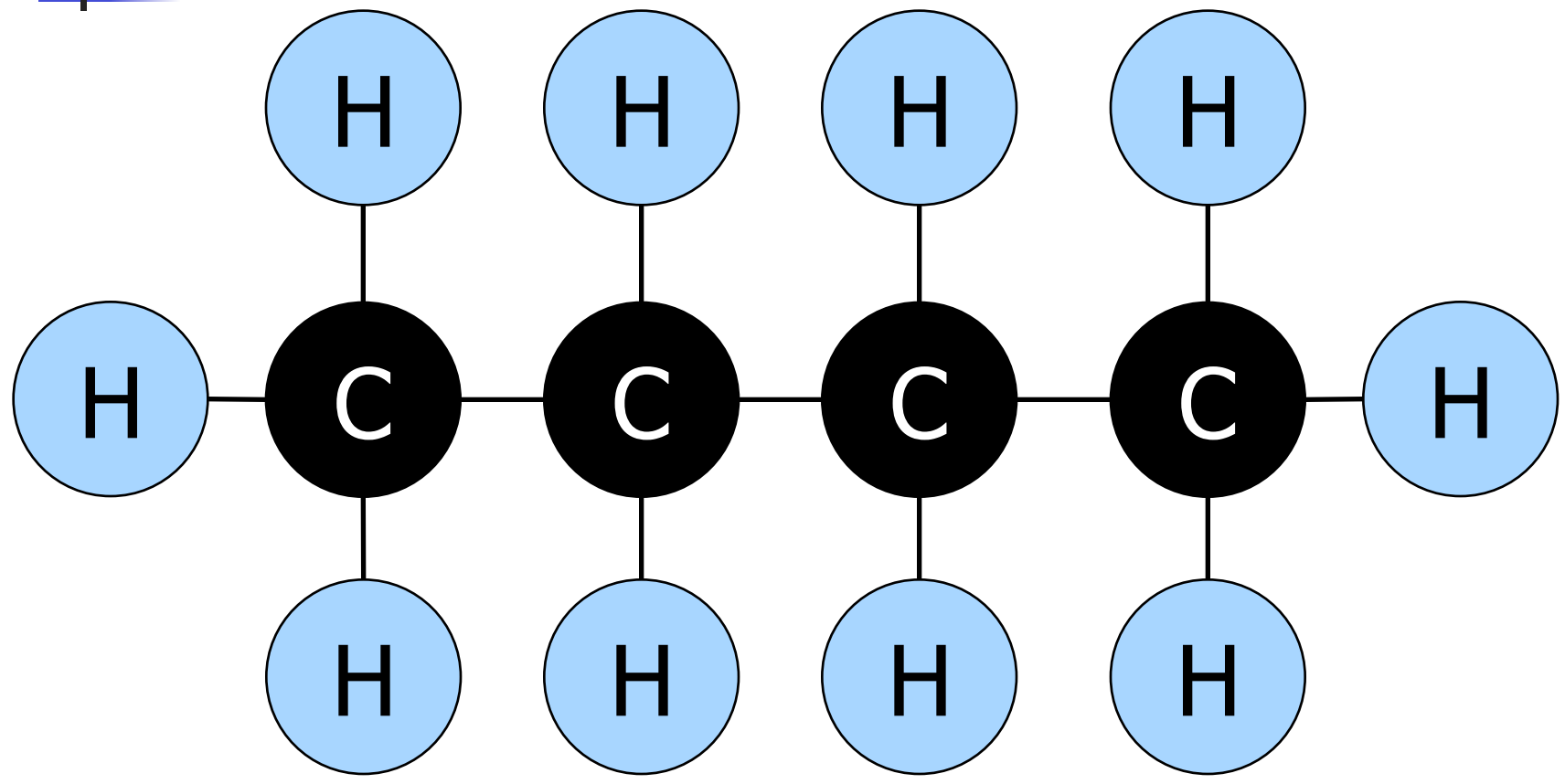


Propane C_3H_8



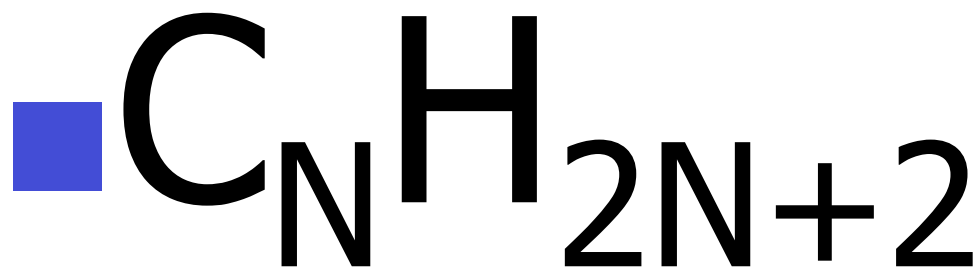


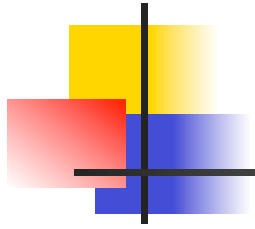
Butane C_4H_{10}



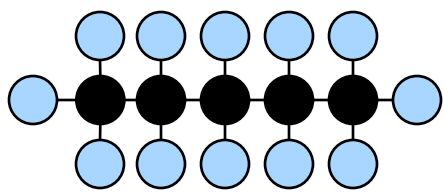


Basic Paraffin Series Formula

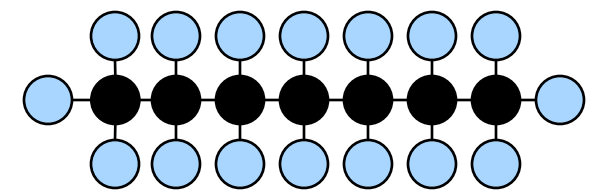




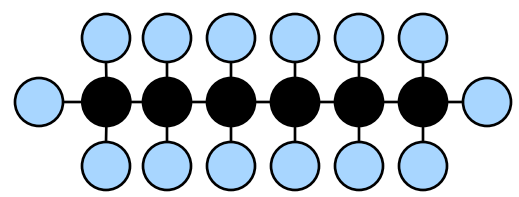
Liquids



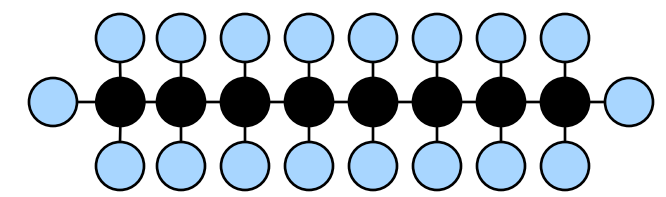
Pentane



Heptane



Hexane



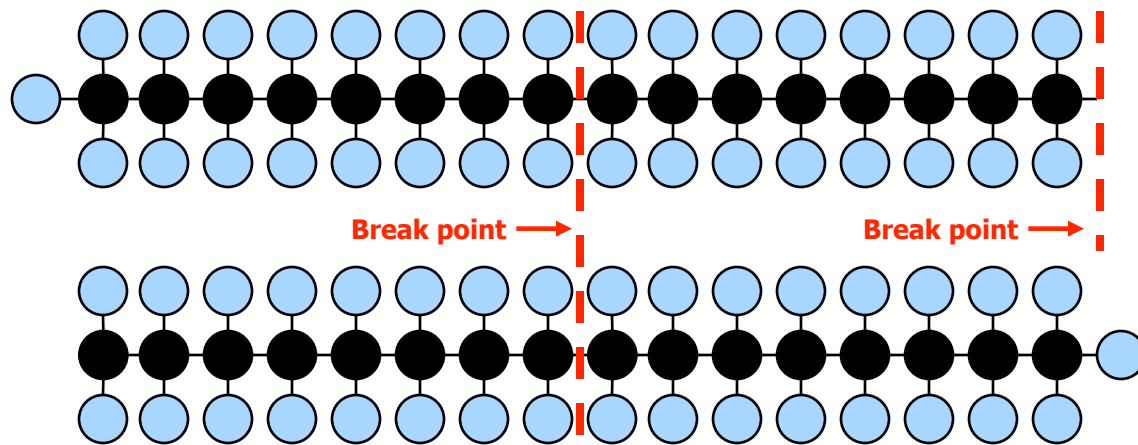
Octane

These hydrocarbons continue to be liquids through C=17



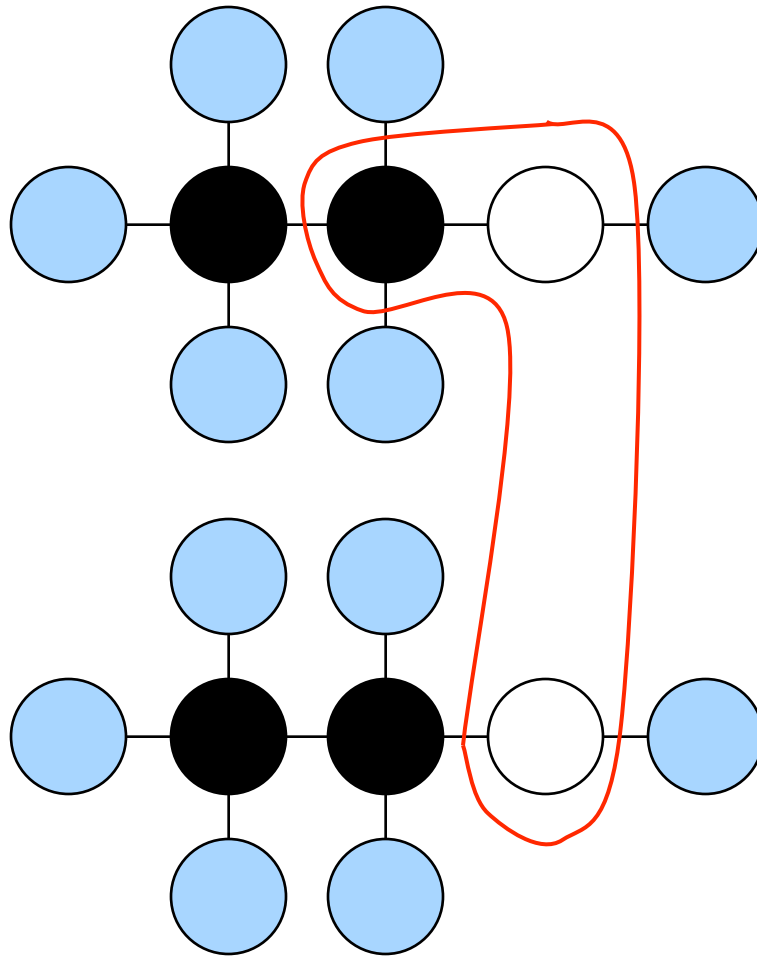
Kerogen

- Any molecule with 18 or more carbons is Kerogen

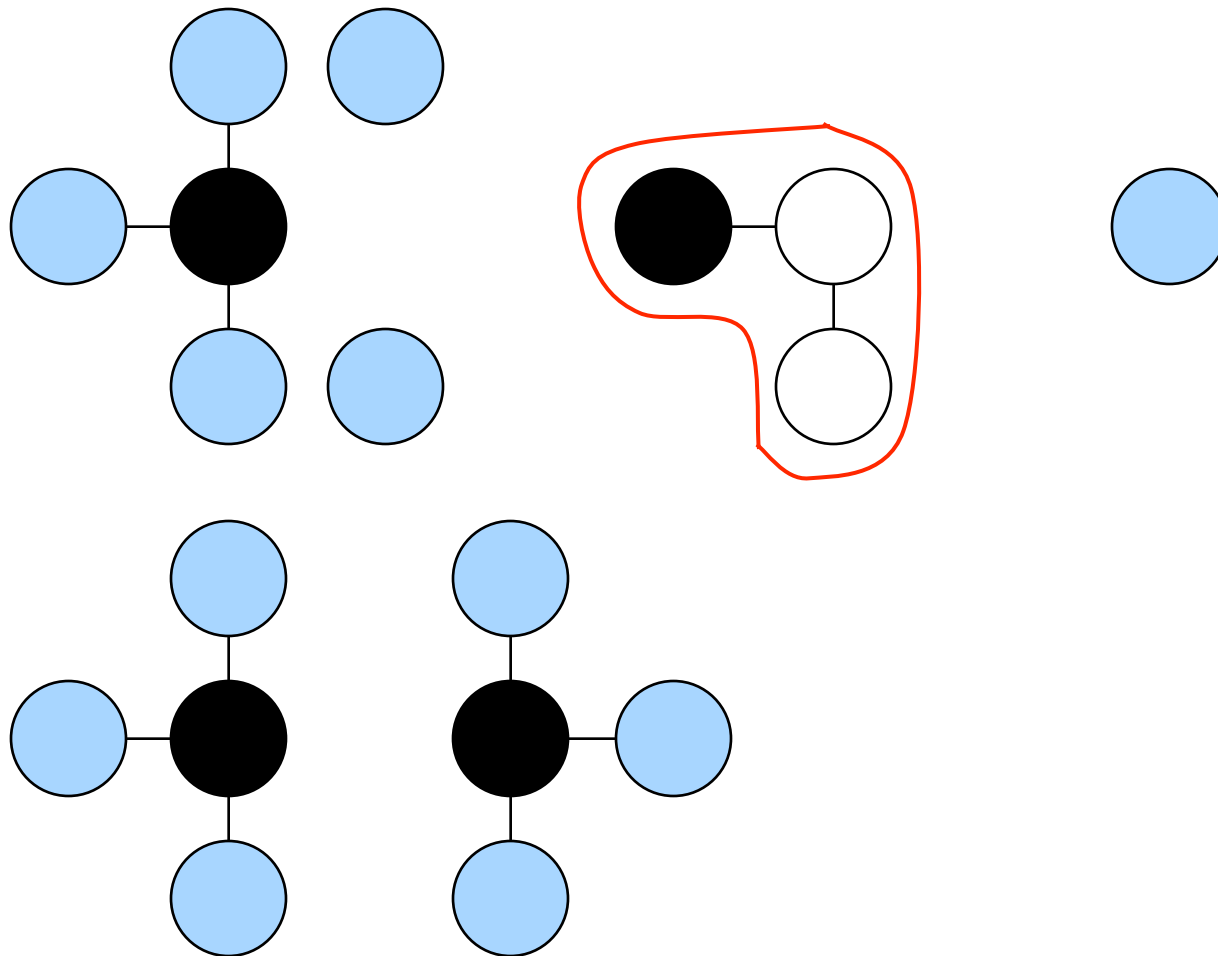


One of the infinitely possible molecules comprising Kerogen

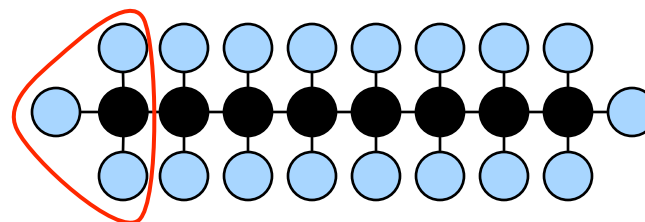
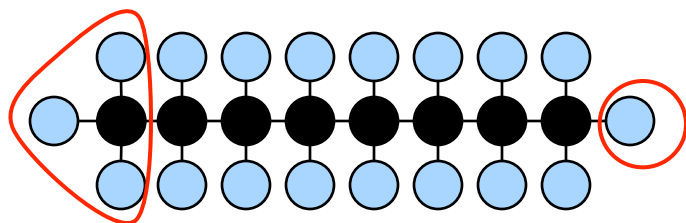
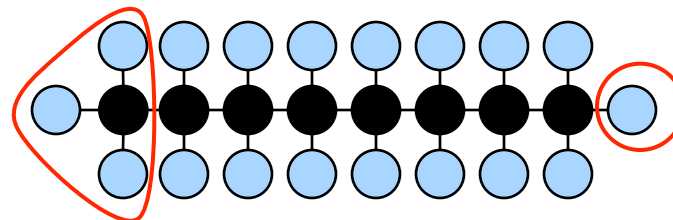
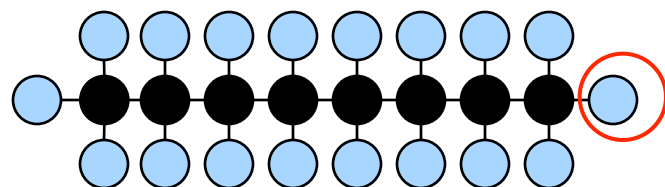
Two Ethyl Alcohol Molecules



Two Ethyl Alcohol Molecules Fractured by Heat



Breaking $C_{32}H_{66}$ into Four Segments



Each segment must have a hydrogen atom at both ends.



Conclusion

- The government survey states that the very best oil shale deposits will yield 25 gallons per ton.¹
- 25 gallons weigh approximately 150 pounds.
 - This is a 7.5% conversion.
- This patented process yields 105% because the alcohol adds to the weight of the synthetic petroleum.
- Therefore, there are 9.4 times more synthetic petroleum available than previously thought.

¹http://pubs.usgs.gov/sir/2005/5294/pdf/sir5294_508.pdf



Questions?

- Thank you.



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